# Calculating pH from Strong Acids

Questions #1-5:  Calculate the pH of each strong acid.

   1)  A 1 x $10^{-3}$ $M$ solution of HCl.

   2)  100. mL of a 1 x $10^{-5}$ $M$ solution of hydrobromic acid.

   3)  2.00 L of a 3.0 x $10^{-4}$ $M$ solution of nitric acid.

   4)  A 0.0525 $M$ solution of hydrochloric acid.

   5)  200. mL of a 0.0100 $M$ solution of HF.

# Calculating [$H^+$] from pH

Questions #6-10:  Calculate the [$H^+$] from each pH value.

   6)  pH = 4.00 of hydrochloric acid.

   7)  HI with a pH of 2.00.

   8)  $HNO_3$ with a pH of 4.50.

   9)  pH = 2.43 of a strong acid.

   10) pH = 5.25 of a weak acid.

# Calculating pH from Strong Bases

Questions #11-15:  Calculate the pH of each strong base.

   11) A 1 x $10^{-4}$ $M$ solution of NaOH.

   12) 250. mL of a 0.0100 $M$ solution of KOH.

   13) pOH = 3.50 for a strong base.

   14) A solution of sodium hydroxide with a [$OH^-$] = 3.40 x $10^{-6}$ $M$.

   15) A strong base with a concentration of 5.83 x $10^{-3}$ $M$.