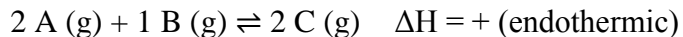


## Final Topics of Equilibrium - Vodcast E-Quiz

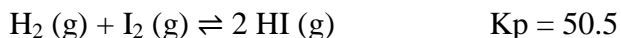
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Questions 1-4 are based upon the following reaction:



According to LeChatlier's Principle, will the reaction shift to the left or right (reactants or products) if the following is done to a system at equilibrium.

- 1) Increasing temperature
    - (A) Shift toward the left (reactants)
    - (B) Shift toward the right (products)
    - (C) No shift
  
  - 2) Increase in pressure
    - (A) Shift toward the left (reactants)
    - (B) Shift toward the right (products)
    - (C) No shift
  
  - 3) Increase moles of A
    - (A) Shift toward the left (reactants)
    - (B) Shift toward the right (products)
    - (C) No shift
  
  - 4) Add a catalyst
    - (A) Shift toward the left (reactants)
    - (B) Shift toward the right (products)
    - (C) No shift
- 



- 5) If 1 atm of  $\text{H}_2$  (g),  $\text{I}_2$  (g), and HI (g) are placed in a sealed flask at a constant temperature, what will occur?
  - (A) The pressure of HI will increase
  - (B) The pressures of  $\text{H}_2$  and  $\text{I}_2$  will increase
  - (C) The pressures will stay constant